

ANSWERS TO TOPIC 1

Review Questions

1. 1 2. 3 3. 2
4. 4 5. 2 6. 2
7. 1 8. 3 9. 3
10. The plum pudding model pictures the atom with an even distribution of positive charges, with electrons throughout the atom. The planetary model pictures the nucleus at the center of the atom with electrons in orbits.
11. Rutherford fired positively charged alpha particles at gold foil. He expected them to pass through with little deflection. When some were greatly deflected, he reasoned that they must have been repelled by a massive positively charged particle, which was called the nucleus.
12. The Rutherford model pictured the electrons in fixed, orbiting positions. The wave-mechanical model pictures electrons as possessing both wave and particle properties.
13. 3 14. 3 15. 1
16. 3 17. 4 18. 4
19. 2 20. 4 21. 2
22. 1 23. 1 24. 3
25. 3 26. 2 27. 2
28. 3 29. 3 30. 2
31. 3 32. 4 33. 4
34. 4 35. 1 36. 2
37. 1 38. 1 39. 3
40. 1 41. 2 42. 3
43. 2 44. 4 45. 4
46. 3 47. 3 48. 4
49. 2 50. 3 51. 2
52. 2 53. 3 54. 1
55. 3 56. 3
57. Na $1s^2 2s^2 2p^6 3s^1$
Mg $1s^2 2s^2 2p^6 3s^2$
Al $1s^2 2s^2 2p^6 3s^2 3p^1$
Si $1s^2 2s^2 2p^6 3s^2 3p^2$
P $1s^2 2s^2 2p^6 3s^2 3p^3$
S $1s^2 2s^2 2p^6 3s^2 3p^4$
Cl $1s^2 2s^2 2p^6 3s^2 3p^5$
Ar $1s^2 2s^2 2p^6 3s^2 3p^4$
58. 1 59. 4 60. 1
61. 1 62. 2 63. 4
64. 4 65. 2 66. 2
67. 2 68. 2 69. 1
70. The mixture contains at least three parts. One solid remained on the filter paper. The other two parts consist of a liquid with a solid dissolved in it.
71. If the calcium and sulfur can be separated by physical means, the substance is a mixture. If chemical means are required to separate them, it is a compound.
72. The picture should show several of each type of molecule evenly mixed.
73. A mixture of iron and oxygen would have the iron present as a solid and the oxygen present as a gas. The iron would be attracted to a magnet. A glowing splint would ignite into a flame if placed into the oxygen gas. If the iron and oxygen were

combined as a compound, it would be a solid that is not attracted to a magnet.

74. Pepperoni pizza is heterogeneous because individual parts can be seen.
75. A is a mixture of an element and a compound; one molecule contains identical atoms and the other one contains different types of atoms. B is a compound; all particles are identical and are made up of two different types of atoms. C is an element; all molecules are identical and each is made up of two identical atoms. D is a mixture of an element and two different compounds; one molecule contains identical atoms and the other two contain different types of atoms, but they differ from each other.
76. Container D is homogeneous because the particles are mixed on the molecular level.
77. The two substances are different materials because their composition varies.

Regents Practice Questions

Part A

- | | | |
|-------|-------|-------|
| 1. 1 | 2. 1 | 3. 1 |
| 4. 3 | 5. 1 | 6. 4 |
| 7. 2 | 8. 3 | 9. 3 |
| 10. 2 | 11. 3 | 12. 1 |
| 13. 2 | 14. 3 | |

Part B

- | | | |
|-------|-------|-------|
| 15. 4 | 16. 3 | 17. 1 |
| 18. 3 | 19. 1 | 20. 3 |
| 21. 1 | 22. 3 | 23. 2 |
| 24. 4 | | |

Part C

25. Because the two nuclei have different numbers of protons, they have different amounts of attraction for electrons. As a result, the amounts of energy separating the energy levels differ for the two elements.
26. The samples are not the same substance because the percent of each element differs from sample to sample.
27. Because he assumed the atom to have an even distribution of positive charges, he did not expect a concentration of positive charges that would strongly deflect or bounce back the alpha particles.
28. The spectra are based on the elements present in the compound. If some lines matched and some didn't, the compounds have a least one element in common and at least one differing element.
29. Column 1: 1, 2, 1; column 2: K, 39, 19; column 3: 92, 146, 92
30. $(0.9778)(12) + (0.0222)(13) = 12.02$
31. They contain different numbers of neutrons.
32. $(19.99 \times .909) + (20.99 \times 0.03) + (21.99 \times 0.88)$
33. 20
34. $(12.00 \times 9.93) + (13.00 \times 1.07)$
35. Both isotopes contain 6 protons but different numbers of neutrons. Or: An isotope with a mass of 12.00 contains 6 protons and 6 neutrons, while the isotope with a mass of 13.00 contains 6 protons and 7 neutrons.

Living Environment Regents Answer Key

Lewis Morris



Living Environment Regents Answer Key :

Answer Key for Regents Living Environment ,2008-09-01 Answer Key for past New York State Regents in Living Environment The Living Environment Prentice Hall (School Division), From basic cell structures to scientific inquiry and lab skills this brief review guides students through their preparation for The Living Environment Regents Examination The book is organized into nine topics each covering a major area of the curriculum and includes a recap of core content as well as review and practice questions vocabulary and six recent Regents Examinations *Regents Exams and Answers: Living Environment, Fourth Edition* Gregory Scott Hunter,2024-01-02 Be prepared for exam day with Barron s Trusted content from experts Barron s Regents Exams and Answers Living Environment provides essential review for students taking the Living Environment Regents and includes actual exams administered for the course thorough answer explanations and overview of the exam This edition features Four actual Regents exams to help students get familiar with the test format Review questions grouped by topic to help refresh skills learned in class Thorough answer explanations for all questions Score analysis charts to help identify strengths and weaknesses Study tips and test taking strategies The Living Environment: Prentice Hall Br John Bartsch,2009 Living Environment Vocabulary Workbook Lewis Morris, Learn the Secret to Success on the Living Environment Regents Exam Ever wonder why learning comes so easily to some people This remarkable workbook reveals a system that shows you how to learn faster easier and without frustration By mastering the hidden language of the subject and exams you will be poised to tackle the toughest of questions with ease We ve discovered that the key to success on the Living Environment Regents Exam lies with mastering the Insider s Language of the subject People who score high on their exams have a strong working vocabulary in the subject tested They know how to decode the vocabulary of the subject and use this as a model for test success People with a strong Insider s Language consistently Perform better on their Exams Learn faster and retain more information Feel more confident in their courses Perform better in upper level courses Gain more satisfaction in learning The Living Environment Regents Exam Vocabulary Workbook is different from traditional review books because it focuses on the exam s Insider s Language It is an outstanding supplement to a traditional review program It helps your preparation for the exam become easier and more efficient The strategies puzzles and questions give you enough exposure to the Insider Language to use it with confidence and make it part of your long term memory The Living Environment Regents Exam Vocabulary Workbook is an awesome tool to use before a course of study as it will help you develop a strong working Insider s Language before you even begin your review Learn the Secret to Success After nearly 20 years of teaching Lewis Morris discovered a startling fact Most students didn t struggle with the subject they struggled with the language It was never about brains or ability His students simply didn t have the knowledge of the specific language needed to succeed Through experimentation and research he discovered that for any subject there was a list of essential words that when mastered unlocked a student s ability to progress in the subject Lewis called this set of vocabulary the

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Environment Lorraine Godlewski,2001-06-30 **High Marks** Sharon H. Welcher,2009-10 *Let's Review Regents: Living*

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UPCO's Living Environment - BIOLOGY Lorraine Godlewski, 2010-09

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